Ka values

A) Using concentrations between 5 x 10^{-3} M to 5 x 10^{-2} M solution for the acid you have chosen calculate the Ka value using the pH probe.

** NB: If the Ka you want to measure is a Ka_2 or Ka_3 then you must use the appropriate salt to make up the solution. Use the molar masses on the bottle as some acids and salts are hydrated.

* we are finding the Ka ( ) for …… using ……..…….

Actual Ka:.....

Percent Error =