Ka values
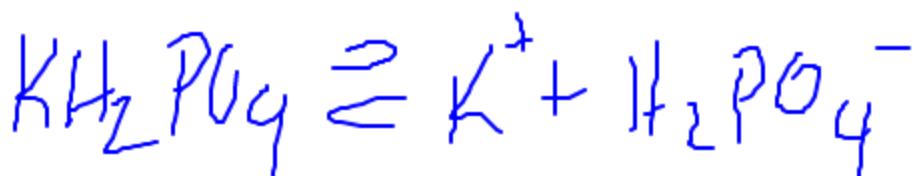
A) Using concentrations between 5×10^{-3} M to 5×10^{-2} M solution for the acid you have chosen calculate the Ka value using the pH probe.

  ** NB: If the Ka you want to measure is a Ka_2 or Ka_3 then you must use the appropriate salt to make up the solution. Use the molar masses on the bottle as some acids and salts are hydrated.

  * we are finding the Ka (   ) for …$H_2PO_4^{-1}$…………… using ……KH_2PO_4………

$$KH_2PO_4 \rightleftharpoons K^+ + H_2PO_4^-$$

$$H_2PO_4^- \rightleftharpoons H^+ + HPO_4^{-2}$$

$$\begin{array}{l} I \\ C \\ E \end{array}$$

Students use the probes to find pH then calculate the Ka from the known values

Actual Ka:………………………………………

Percent Error =